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Photocatalytic Degradation of Commercially Important Dye Acid yellow 36 by $CeFeO_3$

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Abstract

In present work efficiency of $CeFeO_3$as photocatalyst for degradation of Acid yellow 36 was studied.Photocatalytic bleaching of Acidyellow36was observed spectrophotometrically.Theinfluence of variation of different reaction parameters such as concentration of dye; pH and amount ofphoto catalyst were observed.The optimum conditions for photocatalytic degradation have been established.The optimum conditions for the degradation of Acidyellow36dye have been found as dye concentration $3x10^{-5}$M, pH= 8.0 and 0.06 g/100 ml catalyst dose.The kinetics of degradation of the dye follows first order kinetics.A tentative mechanismfor degradation of dye by $CeFeO_3$has been proposed. The presence of hydroxyl radical as an active Oxidising species has been confirmed.

***Key words**: Acid yellow 36,Photocatalyst, Photocatalytic bleaching.*

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I. Introduction

Azo dyes are versatile class of colored organic compounds, which are characterized by the presence of one or more azo bonds (-N=N-). These are widely used in number of industries such as textile dying, food, leather, additive, cosmetic, paper, pharmaceutical industries etc.[1-2].

During dye production and textile manufacturing processes, a large quantity of waste water containing dyestuffs with intensive color and toxicity are introduced into the aquatic systems [3].These dyes do not decompose rapidly through natural processes and are resistant to aerobic degradation. The azo linkage is reduced to aromatic amines under anaerobic conditions that can be toxic and potentially carcinogenic [4-6].

It is necessary to find an effective method of wastewater treatment in order to remove color from effluents. A number of physical and chemical techniques has been reported for the removal of dye compounds such as adsorption on activated carbon [7], biodegradation [8], ozonation [9] and advanced oxidation processes (AOPs) such as Fenton and photo-Fenton catalytic reactions [10,11], H_2O_2/UV processes [12] and semiconductor photocatalysis [13–15].

Advance oxidation processes are the methods in which very reactive speciessuch as hydroxyl radicals ($^{\cdot}OH$)are produced. The hydroxyl radical is a powerful oxidant that can rapidly and non-selectively oxidize organic contaminants into carbon dioxide and water [16,17], so it is able to degrade pollutants effectively[18,19].Among the AOPs heterogeneous photocatalytic oxidation using $CeFeO_3$ as photocatalyst has been extensively studied. $CeFeO_3$is very effective, relatively inexpensive, easily available and chemically stable photocatalyst. The appropriate illumination of these particles produces excited-state high energetic electron and hole pairs (e^-/h^+). These pairs are able to initiate a wide range of chemical reactions that may lead to complete mineralization of organic and inorganic pollutants [20-22].

II. Materials And Methods

All reagents of analytic grade ware used as received without further purification. The $CeFeO_3$ nanoparticles were synthesized by co-precipitation followed by calcinations. The synthesized particles were characterized and used as photocatalyst for degradation of an azodyeAcidyellow36Metal nitrates of Cerium and iron were used for synthesis of photocatalyst because Nitrate group act as oxidizing agent and have high solubility in water.

Fig I: Chemical Structure of Acid yellow 36($C_{18}H_{16}N_3NaO_3S$)
Molecular Weight: 377.39 **λ_{max}:** 512 nm

III. METHOD

The cerium iron oxide particles were synthesized by precipitation of metal nitrates of cerium and iron. The $Ce(N0_3)_3.6H_20$ and $Fe(N0_3)_3.9H_20$ were respectively dissolved in distilled water at room temperature to produce 0.1 M aqueous solution. The ternary oxide $CeFeO_3$waspreparedby adding l M NaOH drop wise to the solution of $Ce(N0_3)_3.6H_20$ and $Fe(N0_3)3.9H_20$ with continuousstirring. The pH of solution was monitored and complete precipitation was obtain at pH 9.The precipitate was filtered and washed. This precipitate was kept in oven at 120° C for drying. After drying it was grounded in mortar and pastles. Subsequently, calcinations were performed. The powder was calcined at 600° C for 5 hours.

Characterization:

Power XRD was performed with Rigaku Micromax-OO7 HF instrument using Cu Ka (ʎ=l.54A°) radiation. The stoichiometry of The CeFeO3 has been established with the help of published diffraction data file JCPDS. All peaks can be indexed to the typical perovskii structure. (JCPDS card 22-0166).

Experimental Procedure:

The degradation of was studied in presence of photocatalyst $CeFeO_3$ at different pH level, catalyst loading and dye concentration. 1X 10^{-3} M dye solution was prepared by dissolving 0.461 g of dye in 1000 ml of distilled water. The initial absorbance of dye solutionwas observed with the help of UV-VIS spectrophotometer(Shimadzu, UV-700, Pharmaspec). The maximum absorbance Value 0.652 was recorded at 512 nm ($ʎ_{max}$). The reaction mixture was prepared by taking 3 ml of dye solution ($1x10^{-3}$ M), 0.06 g of Cerium iron oxide in a round bottom flask. The total volume of the reaction mixture was made 100 ml by adding double distilled water. The concentration of dye in the reaction mixture was $3.0x10^{-5}$ M. To carry out the photobleaching, the reaction mixture was irradiated under light source (2 x 200 W Tungsten lamps). Water filters were used to cut off thermal radiation. The pH of the solution was measuredby pH meter (Systronics, 106). The progress of the reaction was observed at definite time intervals by measuring absorbance using spectrophotometer at 620 nm. The rate of decrease of colour with time was continuously monitored. After complete mineralization, the presence of $N0_2^-$, $N0_3^-$ ions and evolution of CO_2 were tested by standardprocedures. The pH of the reaction was adjusted by adding 0.1 M NaOH and 0.1M HClsolutions.

IV. RESULTS AND DISCUSSION

The synthesized $CeFeO_3$ Was analyzed for its composition and crystal structure by powder XRD. The XRD pattern of $CeFeO_3$ is shown in figure II. The graph has been plotted between intensity and 20 value (in degrees). All peaks can be indexed to typical perovskite structure (JCPDS Card 22-0166). Any peak of impurity was not observed. The sharp and narrow peaks indicates high crystallization of prepared $CeFeO_3$.

Fig II XRD synthesized $CeFeO_3$

The result of photocatalytic bleaching of Acid yellow 36 are given in table 1 .
Table: 1 TYPICAL RUN

[Acid yellow 36] = 3×10^{-5} M Amount of catalyst 0.06 g/100 ml
λ_{max} = 512 nm pH = 8.0

A typical run of photocatalytic degradation of Acid Yellow 36

Time (min)	Abs	1+log abs
0	0.652	0.8142
20	0.600	0.7782
40	0.555	0.7443
60	0.516	0.7126
80	0.477	0.6785
100	0.441	0.6444
120	0.410	0.6128
140	0.378	0.5775
160	0.352	0.5465

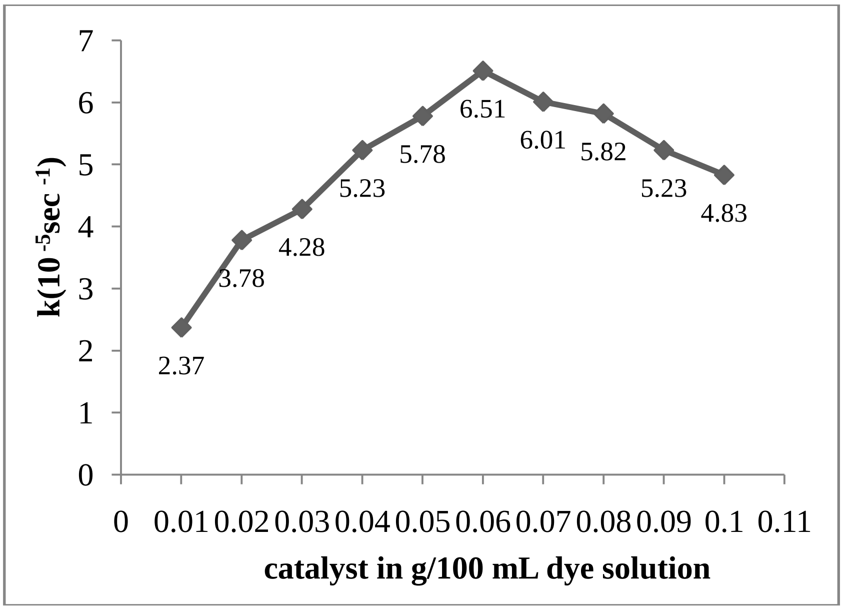
It was observed that absorbence decreases with the increase in time of irradiation indicating that the dye is on degraded irradiation in presence of photocatalyst $CeFe0_3$. A graph between 1 + log abs and time has been plotted. The linearity of the plot indicates that the photocatalytic bleaching ofAcid yellow 36follows a first order kinetics. The rate constant of this photobleaching process was determined using the expression.

Rate constant (k)=2.303 x Slope = 6.51×10^{-5} sec^{-1}

The effect of variation in various reaction parameters has been studied e.g. pH, concentration of the dye, amount of photocatalyst. Control experiments (in absence of photocatalyst, light) confirm the necessity of photocatalyst and light for the photobleaching of dye.

Effect of amount of catalyst on the rate of decolorization of the dye:
Keeping all other factors identical, the concentration of catalyst was changed and its effect on the rate of photochemical degradation was observed.

Table II : Effect of amount of catalyst on rate of decolorization of Acid yellow 36



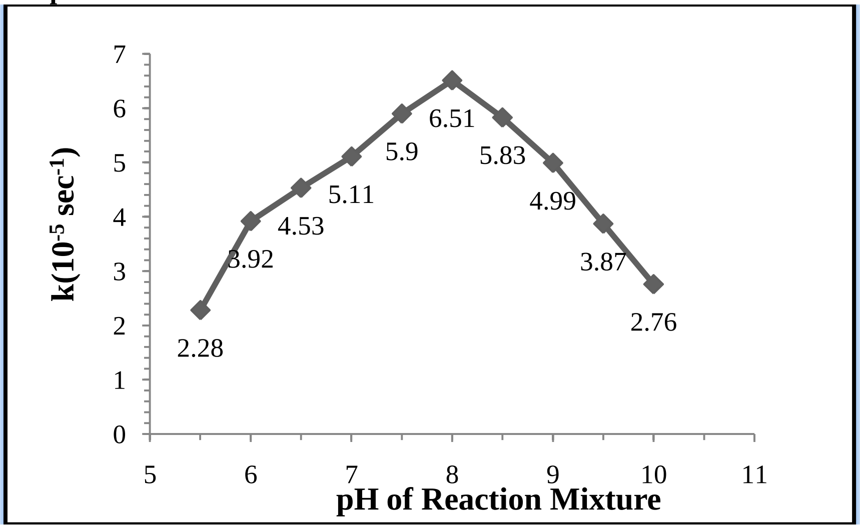
A graph showing effect of variation in amount catalyst on photodegradation of AY 36 by CeFeO$_3$

The result reveals that the rate of photobleaching of dye increases with the increase in the amount of catalyst CeFe0$_3$ up to 0.06g/ 100 ml of dye solution and beyond this the rate of reaction becomes almost constant. The increase in amount of catalyst in the reaction mixture isaccompanied by enhanced generation of OH$^\cdot$ radicals as the surface area of catalyst increase OH$^\cdot$ radical is an extremely strong oxidant which degrade the dye molecule. Furthure increase in photocatalyst above 0.05 g has negligible effect on dye degradation. The amount of catalyst increase with same concentration of dye, further dye molecules are not available for adsorption. Therefore additional catalyst particles are not involved in catalytic activity and rate of reaction is constant.

Effect of variation in hydrogen ion concentration on the rate of decolorization of the dye:

In this study, photodegradation was performed at different pH from 4.5 to 11.0. It has beenobserved that pH 8 is the optimal pH for degradation of Acid yellow 36 Hence all subsequent experiments were carried out at pH 8. After pH 8 there is an adverse effect on the rate of reaction increasing pH further. The observation can be explained on the basis that as pH of solution increases more 0H ions are available. These OH will generate OH radicals by combining withthe hole semiconductor photocatalyst. The pH of reaction mixture has a significant effect on the surface properties of CeFeO$_3$ catalyst hence photocatalytic activity.

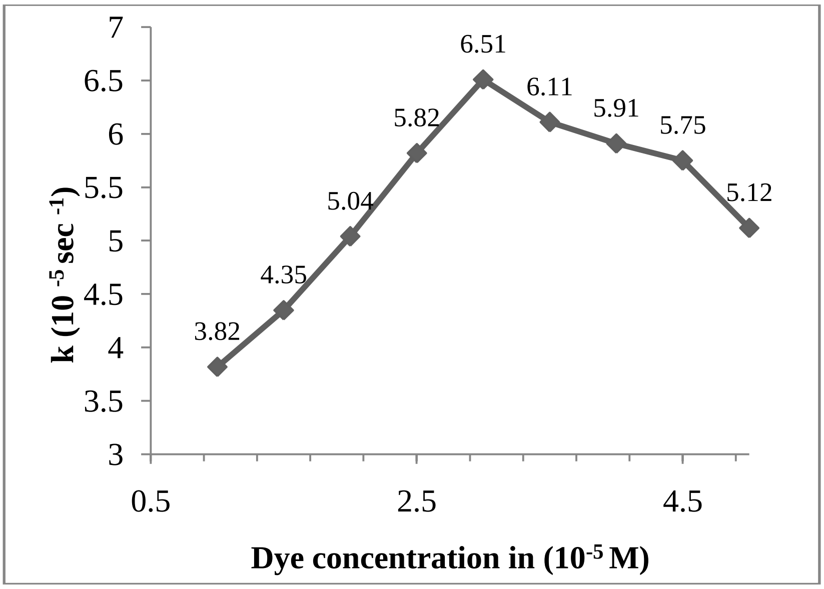
Table III: Effect of pH on rate of decolorization of AY 36



Effect of variation in dye concentration on the rate of decolorization of the dye:

The Effect of dye concentration on the degradation of Acidyellow36 was studied at different concentrations varying from $1.0x10^{-5}$ to $5.5x10^{-5}$ keeping all other factors identical. The result reveals that initially rate of reaction increase with Increase in concentration of dye. The highest rate of reaction was Observed for $3.0x10^{-5}$ M dye solution. Further the rate of photobleaching ofdye decreases with the increase in the concentration of the dye. The reason behind it is that the increase in the initial concentration of the dye lies in consistency of the hydroxyl radicals concentration for all the dye molecules and therefore the rate of decolorization increase. Dye molecules adsorb on Catalyst surface and degradation occurs. On increasing the concentration ofdye, keeping catalyst dose constant, catalyst surface gets saturated. Simultaneously intensecolour of dye doesnot permits light to reach photocatalyst. As result rate of degradation decrease.

Table IV: Effect of dye concentration on rate of decolorizationofAcidyellow 36



Mechanism:

Photocatalysis is initiated by absorption of photon with energy equal or greater than the band gap of semiconductor $CeFeO_3$. $CeFeO_3$ absorbs light in visible region. Electrons from valence band shift into conduction band. Thus a hole (h^+) is generated in valence band. The photogenerated hole is strong oxidizing agent and electron in conduction band acts as reducing agent.

$CeFeO_3 + h\nu \rightarrow CeFeO_3^*$

$h^+(vb) + e^-(cb)$

Theseelectronholepairsmigratetocatalystsurfacewhereradicalsaregenerated.

$h^+ + OH^- \rightarrow {}^\bullet OH$

$h^+ + H_2O \rightarrow OH^\bullet + H^+$

$e^- + O_{2(ads)} \rightarrow O_2^{-\bullet}$

$O_2^{-\bullet} + H^+ \rightarrow HO_2^\bullet$

$OH^\bullet$(hydroxylradical)O_2^-

$^\bullet$(superoxideradical)$HO_2^\bullet$(peroxideradical)arestrongoxidizingspeciesandreactwithdyemoleculestooxidizethem.Simultaneouslyadyemoleculeabsorbradiationofsuitableenergyandgetexcitedto itsfirstsingletstatefollowedbyintersystemcrossingtotripletstate.

$^1Dye_0 + h\nu \rightarrow {}^1Dye_1$(singlestate)

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$^1Dye_1 \rightarrow {}^3Dye_1$(triplestate)

$^3Dye_0$+ radicals→colourless end products

TheparticipationofOH$^\bullet$wasconfirmedusingscavenger2-propanol.Therateofreactionwasdrasticallyreducedin thepresenceof2-propanol.TheformationofCO_2,NO_3^-,NO_2^-indegradedreactionmixtureshowsthereisa completemineralizationinthisprocess.Theendproductsareharmlesstotheenvironment.

V. Conclusion

Cerium iron oxide is effectively utilized as photocatalyst for degradation of textile dyeAcidyellow36in aqueous solution. It is found that degradation is dependent on various reaction parameters web as pH, dye concentration, amount ofcatalyst. The optimum condition for degradation is at pH 8, 0.06 g $CeFeO_3$/ 100 ml dye solution and $3x10^{-5}$ M dye concentration. Overall findings suggest need to exploit this photocatalyst for wastewater treatment of textile effluent.

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References:

[1]. Marchena, C.L., Frenzel, R,A., Gomez, S., Pierella, L.B., and Pizzieo, L.R., Appl. Catal. B: Environ.,13-131,187-196, **(2014).**
[2]. Gao, S., Cao, R., Lu, J., Li ,G., Y, Li and H., J. Mater. Chem., 19, 4157-4163, **(2019).**
[3]. Liu, B., Yu, Z.T., Hang, J., Hua, W., Liu, Y,Y., and Ma, Z.F., Inorg. Chem., 50 (18), 8967-8972, **(2015).**
[4]. Zhang, P.P., Peng, J., Pang, H.J., Sha, J.Q., Zhu, M., Wang, D.D., Liu, M.G., and Su, Z.M.,, Crystal Growth Des., 1(7), 2736-2742, **(2011).**
[5]. Fu, H., Li ,Y., Lu ,Y., Chen, W., Wu, Q., Meng, J., Wang, X., Zhang, Z., and Wang, E., Cryst. Growth. Des. 11(2), 458-465, (2014).
[6]. Takashima, T., Nakamura, R., and Hashimoto, K., J. Phys. Chem. C, 113(39), 17247- 17253, **(2009).**
[7]. Tangestaninejad,S., Moghadam, M., Mirkhani, V., Baltork, I.M., and Salavati, H., J. Iran. Chem. Soc., **7**, S161-S174, **(2019).**
[8]. Morris, A.M., Anderson, O.P., and Finke,R.G.,Inorg. Chem., 48 (10), 4411–4420, **(2019).**
[9]. Dey, K.C., Sharma, V., Int. J. Chem. Tech. Res., 2 (1), 368, **(2010).**
[10]. Yajun, W., Kecheng, L.U., and Changgen, F., J. Rare Earths, 29(9), 866-871, **(2016).**
[11]. Alnuaimi, M.M., Rauf, M.A., and Ashraf, S.S., Dyes Pigments, 76 (2), 332-337, **(2018).**
[12]. Peng, H.J., Jang, C.J., Huang, N.N., Huang, D.Y., and Chen, Y.G., Solid State Sci., 35(33-38), **(2014).**
[13]. Ghorai, T.K., Biswas, S.K., and Pramari, P., Appl. Surf. Sci., 254 (22), 7498-7504 **(2018).**
[14]. Yang, S.J., and Tang, J.Y., China Chemicals, 5, 36-38, **(2018).**
[15]. Feng, C., Liu, X., Xu, G., J. Rare Earth, 13(1), 44-48, **(2016).**
[16]. Wang, Y., Lu, K., Feng, C., J. Rare Earth, 31(4), 360-365, **(2017).**
[17]. Zhao, K., Liu, Y., Lu, N., Jhao, Y., Yuan, X., Zhang, X., Tang, L., Li, F., Appl. Surf. Sci., 285(B), 616-624, **(2018).**
[18]. Herrera,JAR.,Frenzel, R.A., Blenco, M.N., Pizzio, L.R., J. Photochem. andPhotobiol.-A, 289, 22-30, **(2014).**
[19]. Barras, A., Cordier, S., Boukherroub, R.,AppliedCatal. B: Env.,123-124,1-8, **(2020).**
[20]. Chai, H., Wang, L., Xu, L., Wang, X., Huang, J., Dyes and pigments, 76(1), 113-117, **(2018).**
[21]. Ghalebi, H.R., Aber, S., Karimi, A., J. Mol. Cat. A: Chem., 415, 96-103, **(2022).**
[22]. Liu, C.G., Zheng, T.,Zheng, H.Y., J. Mol. Struc. 1110, 44-52, **(2023).**